مدرسة جيه اس اس الخاصة
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Lesson 5 - Class Test

Periodic classification of Elements. Date: 01/11/20

1. (i) ~~Groups~~ Group 17

(ii) Group ~~8~~ 17

(iii) Group 1

(iv) Group ~~Zero~~ 1 - 16

(v) Group 18, they are noble gases

2.

(i) The atomic size increases down the group as the number of shells increases.

(ii) Metal have the tendency to lose electrons and form a ~~negative~~ positive ion, therefore they are electropositive in nature.

(iii) Cl is smaller than Ar, this is because they are in the same period and along the period, atomic size decreases.

(iv) Li is smaller than F, they are in ~~same~~ the same period and atomic size decreases along the period.

(v) Cl is more electronegative. It has an atomic number of 17 and 7 valence electrons and hence ~~wants~~ tends to gain an electron to complete octet. This is not the case with ~~Flourine~~ F

3(i) Lithium (Li) ~~and~~ and Sodium (Na)

(ii) Beryllium (Be) and Calcium (Ca)

(iii) Neon (Ne) and Argon (Ar)

(iv) ~~Flourine~~ Fluorine (F) and Chlorine (Cl)

(v) Carbon (C)

4.

(i) The atomic size decreases across the period, this is because the number of valence electrons increases which the increases the Effective Nuclear charge between protons and electrons.

(ii) Group 1 has the largest atomic size. This is because the atomic size decreases along the period. Group 1 elements are before group 17 elements in a period.

(iii) Group 17 are smallest in size. This is because these elements come after group 1 elements and atomic size decreases across the group.

(iv) The number of valence electrons increases and hence the Effective Nuclear charge between protons and electrons increases. Due to this strong nuclear attraction, atomic size decreases along the period.