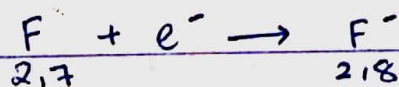
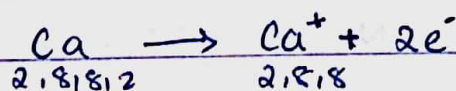
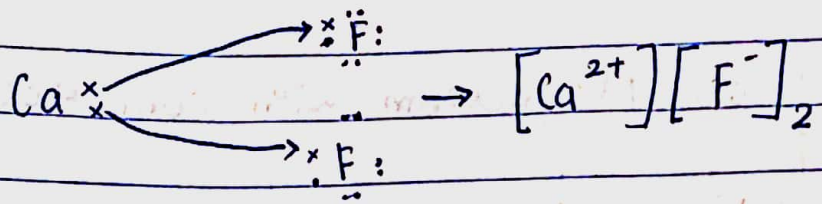


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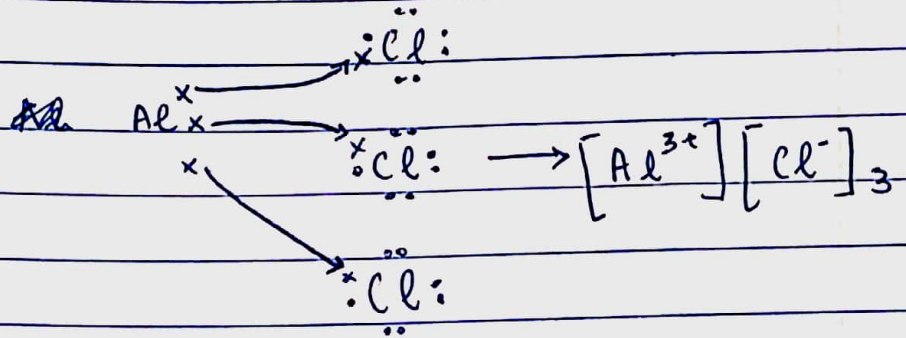
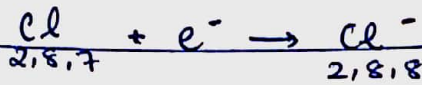
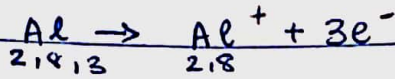
Lesson 3 - Revision Test.

- 1.1. Nitric acid ( $\text{HNO}_3$ ) is a strong oxidising agent. Therefore, it oxidises  $\text{H}_2$  to water and itself gets reduced to nitrates and nitrites.
2. Aluminium oxide ( $\text{Al}_2\text{O}_3$ ) reacts with both acidic and basic substances, to form a salt and water  
With acid  $\text{HCl}$   $\rightarrow$  :  
$$2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2\text{O}$$
With base  $\text{NaOH}$  :  
$$\text{Al} + 2\text{NaOH} \rightarrow \text{NaAlO}_2 + \text{H}_2\text{O}$$
3. In an aqueous solution, the ions dissociate and these free electrons are capable of conducting electricity
4. Metallic oxides, when they react with acidic substances, form ~~salt~~ salt and water. Hence, they are basic
5. Silver articles react with atmospheric ~~oxid~~ oxygen and are corroded over time. Hence, they lose their colour.

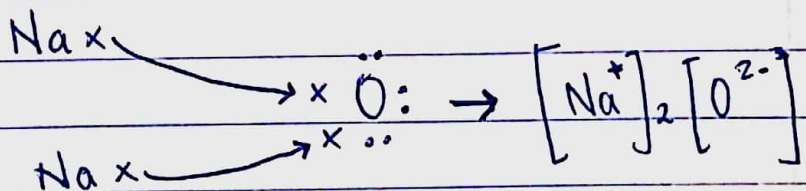
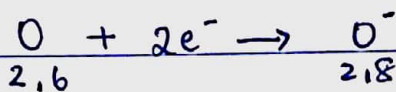
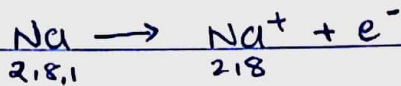
II 6.  $\text{CaF}_2$ 



7.  $\text{AlCl}_3$



8.  $\text{Na}_2\text{O}$



III.9. (d)  $\text{Al} > \text{Cu} > \text{Zn} > \text{Fe}$

10. (c) Zinc is more reactive than tin

11. (c) Magnesium with conc. HCl

12. (d)  $\text{AgNO}_3$  and Cu metal